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 AP Chemistry 1 Text: Section 5.1 (Do not complete the following: 5.1 Activity: Linking the Giants); AP Chemistry 1 Text: Section 5.2 (Complete Activity 5.2; Do not complete the following: 5.2 Review Questions); AP Chemistry 1 Text: Section 5.3 and 5.4 (Do not complete the following: 5.3 Activity: Uncle Werner says, "Consider This" and 5.4 Activity: A Post-it Periodic Table)

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1. If 2.49 g of CuNO<sub>3</sub> is dissolved in water to make a 0.830 M solution, what is the volume of the solution in milliliters? 2. How many moles of NaOH are present in 13.5 mL of 0.170 M NaOH? 3. Calculate the molarity of 0.650 mol of Na<sub>2</sub>S in 1.15 L of solution. 4. A student in lab needs to make a solution that is 7.00% by mass NaCl. If 145 g of NaCl is available, what mass of solution can be ...

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1. A solution made by adding 66.7mL of ethylene glycol to enough water to give 197mL of solution. 2. A student in lab needs to make a solution that is 7.00 % by mass NaCl. If 145 g of NaCl is available, what mass of solution can be prepared?

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